

# Evaluation of Silymarin B-Cyclodextrin Molecular Complexes

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## ABSTRACT

Silymarin (482.43) is not a good water soluble hepatoprotective agent. The oral absorption of silymarin is only about 23 – 47%, leading to low bioavailability of the drug, which limits its use. It is a mixture of mainly three flavanolignans, namely, silybin, silydianin, and silicristine, with silybin being most active. silymarin has been used medicinally to treat liver disorders, including acute and chronic liver hepatitis, toxin –/drug-induced hepatitis and cirrhosis, and alcoholic liver diseases. It is also reported to be effective in certain cancers.[1–5] The poor aqueous solubility of the drug may lead to dissolution-related bioavailability problems. Many approaches such as solubilization with surfactant systems, formation of water soluble complexes, and use of pro-drugs and soluble salt formation have been reported for improving the solubility and dissolution, and in turn the bioavailability of the drug.[6–8] Cyclodextrin and its derivatives play an important role in the formulation development due to their effect on solubility, dissolution rate, chemical stability, and absorption of a drug.[9] Nagarsenkar *et al.* reported faster dissolution and better bioavailability of ketorolac solid dispersion with HP-β-CD.[10] Reddy *et al.* reported enhanced solubility and dissolution rate of Celecoxib by complexation with β-Cyclodextrin.[11] The objective of the present study is to investigate the possibility of improving the solubility and dissolution rate of silymarin by complexation with β-cyclodextrin and also to compare the different complexation methods with respect to their dissolution study. In addition, the physicochemical characteristics of solid inclusion complexes were also investigated. A gift sample of silymarin was obtained from AGAPE pharmaceuticals, Majhitar, East Sikkim; β-cyclodextrin was obtained from Zim Laboratories Ltd., Nagpur. HCl, methanol potassium dihydrogen phosphate, and sodium hydroxide were of analytical grade (Merck, Mumbai). All other chemicals and reagents used were of analytical grade.

**Key Words:** Silymarin, Molecular complexes, Cyclodextrin, evaluation, method, solution

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## INTRODUCTION

### Phase solubility analysis for silymarin

Phase solubility studies were performed to determine the stoichiometric proportions of silymarin with β-cyclodextrin.[12,13] The data was used to determine the stability constant of the complexes. For this, the stock solution of 0.01 M β-cyclodextrin was prepared using distilled water. These stock solutions were diluted with distilled water to give molar solutions in the range of 0.002 to 0.01 M β-cyclodextrin. Five ml of each molar solution was filled in screw capped vials and the excess quantity of the drug was added to each vial separately.[7,14] The vials were shaken at an ambient temperature, for 48 hours, using a laboratory shaker (Remi). The supernatant solutions were collected carefully and filtered using Whatman filter paper (No. 41). The concentration of the drug in filtered solutions was determined using a UV visible spectrophotometer. No changes in λ max of the drug were found after complexation with cyclodextrin, hence absorbance of the resultant solutions were recorded at 286 nm, which was λ max of the drug. From the slope and intercept value (S<sub>0</sub>) of the phase solubility curve, a stability constant (K<sub>c</sub>) was determined.

$$K_c = \text{Slope} / [S_0 (1 - \text{Slope})]$$

Preparation of physical mixture and inclusion complexes

**Physical mixture method**

The required molar (1 : 1) quantities of the drug and cyclodextrin were weighted accurately and mixed together thoroughly in a mortar, with vigorous trituration, for about three hours. These mixtures were then passed through sieve No. 44, and finally stored in airtight containers till further use.[7,15,16]

**Kneading method**

The required quantities of the drug (Silymarin) and  $\beta$ -cyclodextrin were weighed accurately in a ratio of 1:1. A homogenous paste of cyclodextrin was prepared in a mortar by adding water : Methanol mixture (1 : 1) in small quantities. Silymarin powder was then added to this paste in portions, with continuous kneading, for three hours. An appropriate quantity of water : Methanol mixture (1:1) was added further to maintain suitable consistency of the paste. This paste was dried in a hot air oven at 45°–50° for 24 hours. The dried complexes were then powdered and passed through sieve No. 44 and stored in airtight containers till further use.[15]

**Co-precipitation method**

Quantities of drug and cyclodextrin, in the required molar ratio (1 : 1), were dissolved in methanol : Water, respectively. The solution of the drug was added dropwise into cyclodextrin solution. The contents were continuously stirred for 6 hours and finally were dried at 45°–50° for 48 hours, collected, and stored in airtight containers till further use.[16]

**Solvent evaporation method**

In this technique, silymarin along with solubilizing additives such as acetone were dissolved at 25°C temperature. Next, the required moles of  $\beta$ -cyclodextrin in hot distilled water were added dropwise into this solution, with continuous stirring, for one hour. The complexes formed were filtered and dried under a vacuum. Then the prepared solid mass was stored in a desiccator under vacuum to a constant weight. The dried products were removed, pulverized, and passed through sieve No. 100 and finally stored in a closed airtight container.

**Characterization of silymarin inclusion complexes****Drug content estimation**

The quantities of inclusion complex equivalent to 70 mg of silymarin were dissolved in water: Methanol mixture (1:1). Appropriate dilutions were made and the drug content of each complex was calculated from UV absorbance recorded at  $\lambda$  max 286 nm.

**Scanning electron microscopy**

The morphology of the inclusion complexes by physical mixture, kneading method, and co-precipitation method was studied using a scanning electron microscope (JSM-5610 LV Jeol, Japan). The samples were coated with platinum to provide a conductive layer for observing images at 15 kV.

**IR spectrum analysis**

Infra red (IR) spectra of the drug and inclusion complexes were recorded using the KBr method using Fourier Transform Infrared Spectrophotometer (FT IR-8400 S). A baseline correction was made using dried potassium bromide, and then the spectra of the dried mixtures of drug and inclusion complexes with potassium bromide were recorded.

**Differential scanning calorimetric analysis**

This scanning was performed using DSC model (Perkin Elmer). The samples were placed in a closed platinum crucible and DSC thermograms were recorded at a heating rate of 10°/minute in the range of 20° to 310°C.[16]

**X-ray diffraction study**

The X-ray diffraction pattern of the selected inclusion complex was compared with that of pure silymarin. This was performed by measuring  $2\theta$  in the range of 4° to 50° with reproducibility of  $\pm 0^\circ-001^\circ$  on an X-ray diffractometer (Phillips).

**Dissolution study of silymarin and its inclusion complexes**

Dissolution of inclusion complexes (equivalent to 70 mg of silymarin) was studied using the USP XXII eight station dissolution apparatus (Electrolab TDT-08L). The dissolution was carried out in 900 ml of Phosphate buffer, pH 6.8, at a speed of 75 rpm. Aliquots of 10 mL were withdrawn periodically and replaced with 10 mL of fresh dissolution medium. The concentrations of the drug in the samples were determined by measuring their absorbance at 286 nm, using Shimadzu 1700 UV-visible spectrophotometer. Cumulative percent of the drug released was determined at every point of time. The pure drug was used as a control. The  $T_{90}$  (time required for 90% dissolution of drug) of various solid dispersions were calculated.[17,18]

### Phase solubility analysis of silymarin

The phase solubility study was done to determine the stoichiometric proportion of silymarin with complexing agent  $\beta$ -cyclodextrin. The solubility analysis indicated the formation of a 1 : 1 molar inclusion complex of the drug with  $\beta$ -cyclodextrin. The apparent stability constant for silymarin ( $K_c$ ) was  $722 \text{ K}^{-1}$  with the  $\beta$ -cyclodextrin complex.

### Characterization of silymarin inclusion complexes

All the inclusion complexes prepared using different methods, such as, physical mixture, kneading method, co-precipitation, and solvent evaporation method were found to be slightly brown, free-flowing powders.

### Estimation of drug content

Inclusion complexes prepared by co-precipitation showed nearly 100% drug content. The drug content of the inclusion complex prepared by kneading, physical mixture, and solvent evaporation shows slightly less drug content as compared to that prepared by using co-precipitation [Table 1].

**Table 1: Comparison of drug content and T90 value of silymarin inclusion complex with  $\beta$ -cyclodextrin**

Systems	% Drug content	T <sub>90</sub> (min)
Silymarin (alone)	-	<120
Silymarin : $\beta$ cyclodextrin (PM)	97.140	62
Silymarin : $\beta$ cyclodextrin (KN)	91.420	47
Silymarin : $\beta$ cyclodextrin (CP)	99.995	94
Silymarin : $\beta$ cyclodextrin (SE)	97.135	92

PM: Physical mixture, KN: Kneading, CP: Co-precipitation, SE: Solvent evaporation

### Scanning electron microscopy

Scanning electron microscopy (SEM) of the physical mixture method, inclusion complex by kneading method, and co-precipitation method were studied. The representative photographs are shown in Figures Figures11–3. Pure drug particles in the physical mixture method were very small in size with reduced effective surface area, due to agglomeration. They remained dispersed and physically adsorbed on the surface of  $\beta$ -cyclodextrin. Figures Figures22 and and3,3, show drug particles distributed on the surface of  $\beta$ -cyclodextrin. Both kneaded and the co-precipitated systems show a homogeneity, signifying the inclusion complex formation. The kneaded system is of poor crystal structure, lacks distinct crystal faces, and has numerous cracks and fissures. This may also have contributed to faster dissolution compared to the co-precipitation system.

### IR spectral analysis

Infra red spectra of the pure drug and inclusion of silymarin with  $\beta$ -cyclodextrin, by different methods, are given in Figure 4 (a - e). The guest molecules within the cyclodextrin cavity show in their peaks or show peaks of less intensity. Basically peaks that lie in the fingerprint, and peaks due to C-O or O-H stretching are affected (shifted or intensity is changed).

### Differential scanning calorimetric analysis

Silymarin,  $\beta$ -cyclodextrin, and the complex prepared by the co-precipitation method were subjected to Differential scanning calorimetric (DSC) analysis. The thermal curves are given in Figure 5. The DSC curve of silymarin shows one characteristic sharp endothermic peak at around  $150^\circ\text{C}$ , indicating the melting point of the drug. The DSC curve shows that the sharp endothermic peak at around  $150^\circ\text{C}$ , which is observed for silymarin, decreases in the inclusion complex (1:1). Furthermore, the wide peak at  $90^\circ\text{C}$ , which is observed for  $\beta$ -cyclodextrin, shifts to  $60^\circ\text{C}$  in the inclusion complex, indicating that the inclusion complex does not contain much residue of silymarin or  $\beta$ -cyclodextrin, thus suggesting that the drug is well dispersed in the  $\beta$ -cyclodextrin cavity.

### X-ray diffraction study

The inclusion complex of the drug prepared with  $\beta$ -cyclodextrin, using the co-precipitation method, which showed a good result overall, was characterized further by an XRD study The X-ray diffraction patterns of pure silymarin, as well as the silymarin- $\beta$ -Cyclodextrin complex obtained by using the co-precipitation method are represented in Figure 6. The peak position (angle of diffraction) is an indication of the amorphous nature of the sample. The diffractogram of pure silymarin

shows some intense peaks, which are indicative of crystallinity. However, in case of silymarin complexed with  $\beta$ -Cyclodextrin diffractogram, it attributes to a new solid phase with low crystallinity, indicating inclusion complex formation (more water soluble). A reduced number of signals, of markedly low intensity, are noticeable in the complex, indicating the greater amorphous nature of the inclusion complex compared to the free molecules.

#### Dissolution study of silymarin and its inclusion complexes

The inclusion complexes of silymarin with  $\beta$ -cyclodextrin produce pronounced enhancement in its dissolution as shown in Table 2 and Figure 7. From Table 1-T<sub>90</sub> value for kneading shows the quickest rate of dissolution as compared to the other inclusion complexes, prepared by other inclusion methods. Inclusion complexes prepared by the co-precipitation methods show their T<sub>90</sub> in about 94 minutes.

**Table 2: Percentage cumulative release of inclusion complexes of silymarin and  $\beta$ -cyclodextrin prepared by different methods**

Time	%CR (Silymarin)	%CR (PM)	%CR (KN) 9	%CR (CP)	%CR (SE)
5	7.57	34.94	31.71	24.055	15.32
15	14.20	49.56	53.52	27.481	39.25
30	20.05	60.82	83.58	59.763	47.01
45	23.45	80.61	99.08	66.686	77.05
60	27.70	97.76	-	69.520	84.23
75	32.79	-	-	87.200	86.41
90	43.67	-	-	93.455	98.80
105	46.34	-	-	95.320	-
120	49.13	-	-	-	-

PM: Physical mixture, KN: Kneading, CP: Co-precipitation, SE: Solvent evaporation

#### CONCLUSION

The phase solubility data suggest a 1 : 1 complex formation with  $\beta$ -cyclodextrin. All inclusion complexes show increase in dissolution than in the drug alone. The inclusion complex of the drug prepared by the co-precipitation method shows the best results overall, in terms of drug content and dissolution profile, for preparation of sustained release formulations. However, it should be noted that this is a result of a preliminary study of  $\beta$ -cyclodextrin inclusion complexation with silymarin.

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